A structurally flexible triazolate-based metal–organic framework featuring coordinatively unsaturated copper(I) sites†

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The novel metal–organic framework CFA-8 (Coordination Framework Augsburg University-8), [Cu2(tqpt)], containing the organic linker H2-tqpt (H2-tqpt = 6,6,14,14-tetramethyl-6,14-dihydroquinocinolino[2,3-b]phenazinebis-triazole), has been synthesized. Reaction of H2-tqpt and anhydrous CuCl2 in N,N-dimethylacetamide (DMA) yields CFA-8 as orange crystals with lenticular shape. This framework shows a reversible breathing effect and is robust upon solvent removal. It has been characterized by single-crystal and powder X-ray diffraction, TGA, IR spectroscopy and gas sorption measurements. CO adsorption isotherms show that Cu(I) sites in this framework are able to bind carbon monoxide forming a weak complex which has been additionally characterized by IR spectroscopy and synchrotron powder X-ray diffraction measurements.

Introduction

With their patent in 1947 Procter and Gamble Ltd reported on the corrosion inhibiting effect of 1H-benzo[d][1,2,3]triazole (btaH) for copper metal.1 The same molecule was later mentioned in some additional patents describing corrosion inhibitors.2,3 Driven by the economic impact of these inhibitors various coordination modes between copper and 1,2,3-benzotriazoles have been proposed.4–13 Cotton and co-workers published the first Cu(I)-bta structure and stated more decisively that this structure contains Cu(I) ions.4 This coordination mode and the monovalent oxidation state of copper was later confirmed by Poling, but no crystal structure has been reported for a Cu(I) complex derived from 1H-benzo[d][1,2,3]triazole.13,14 The first crystal structure of a Cu(I) polymer incorporating the deprotonated ligand of 1H-benzo[d][1,2,3]triazole was reported by Huang and co-workers.13 The connection scheme of this 1D coordination polymer with the formula [Cu2(bta)3] is shown in Fig. 1. The crystal structure is much more complicated than the original model proposed by Cotton et al., which suggested a linear polymeric structure. The asymmetric unit of this structure contains three differently coordinated Cu(I) ions. One of the CuI ions is trigonally-planar coordinated by nitrogen donor atoms of the triazolate ring, the second CuI ion is two-fold linearly coordinated and the third copper ion is tetrahedrally coordinated. The most interesting feature of this compound is the presence of three different coordination modes of the copper(I) ions. With a more focused look at the structure it could be assumed that the two- and three-fold coordinated Cu(I) ions should be able to bind additional ligands by keeping their original connection scheme, which would however require to engineer the dense Cu(I)(bta) 1-D coordination polymer into a porous 3D framework structure.

†Electronic supplementary information (ESI) available: IR spectra, details of gas sorption measurements, UV/vis spectrum and crystallographic data. CCDC 1483639-1483641. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6dt02672a
nation modes, which we refer to as CFA-8 (Coordination Framework Augsburg University-8). This MOF represents a further member of Cu(I)-containing metal–organic frameworks comprising open Cu⁺ metal sites, which might be amenable to binding of additional ligands.¹⁵–¹⁷ To the best of our knowledge, the crystal structure of the framework introduced here is the first example for a crystalline and permanently porous Cu(I) network incorporating a bistriazolate as organic bridging ligand.

Results and discussion

Syntheses and characterization

The organic ligand H₂-tqpt was synthesized in three steps. 2,2,5,5-Tetramethylcyclohexane-1,3-dione and 3,3,6,6-tetramethylcyclohexane-1,2,4,5-tetraone were synthesized by conventional procedures.¹⁸ The condensation reaction of the latter tetraone and two molecules of 1H-benzotriazole-5,6-diamine¹⁹–²¹ leads to the organic bistriazolate ligand H₂-tqpt, which was previously used for the synthesis of the CFA-7 metal–organic framework (Scheme 1).²²

Metal–organic framework CFA-8 can be synthesized from H₂-tqpt and anhydrous CuCl₂ in a good yield by using N,N-dimethylacetamide (DMA) as a solvent at 140 °C (Scheme 2). During the reaction, Cu(II) ions are reduced by DMA to Cu(I). The network forms orange lenticular crystals (Fig. 2). Inspection of electron microscopic images reveals that the crystal faces of the crystals stay intact, even after activation at 250 °C in vacuum. Different microscope pictures show that CFA-8 forms crystals in the size range of 10–150 μm.

Crystal structure analysis

A single crystal X-ray diffraction study reveals that CFA-8 crystallizes in the monoclinic crystal system in the space group C2/c (no. 15), which coincides with the crystal structure described for Cu(I)(bta) which also crystallizes in the monoclinic crystal system, in the same space group.¹³ The asymmetric unit of CFA-8 consists of three copper, twenty-two carbon, sixteen hydrogen and ten nitrogen atoms (Fig. 3).

The three crystallographically independent copper ions have different coordination geometries. The Cu₃ atom (see asymmetric unit Fig. 3) located on a site with crystallographically imposed twofold symmetry is linearly two-fold coordinated by two nitrogen atoms (N₁₀, N₁₀a) of different tqpt⁻ ligands. The distances between Cu₃ and N₁₀, N₁₀a are 1.874(3) Å and the angle N₁₀–Cu₃–N₁₀a is with 178.45(18)° nearly straight (Fig. 4). The trigonally planar coordinated Cu₁, located on a general position, binds to three nitrogen atoms (N₁, N₃ and N₉). The Cu₁–N₁/3/9 bond lengths range from 1.931(3)–1.995(3) Å and the N–Cu₁–N angles fall in between

![Scheme 1](image1.png)  
**Scheme 1** Synthesis of the organic bistriazolate ligand H₂-tqpt.

![Scheme 2](image2.png)  
**Scheme 2** Synthesis conditions to obtain CFA-8.

![Fig. 3](image3.png)  
**Fig. 3** Ball-and-stick model of the asymmetric unit of CFA-8 (grey: C-, blue: N-, cyan: Cu-atoms, H-atoms are omitted for clarity; atoms displayed as 50% probability ellipsoids).

![Fig. 4](image4.png)  
**Fig. 4** Different coordination environments of Cu(I) ions in CFA-8 (CuLa (−1 + x, −1 + y, −1 + z); N1a (−1 + x, −1 + y, −1 + z); N2a (1 − x, −1 + y, 0.5 − z); N3a (−1 + x, − y, −0.5 + z); N8a (−0.5 + x, −0.5 + y, z); N9a (−0.5 + x, −0.5 + y, −1 + z); N10a (−0.5 + x, −0.5 + y, −1 + z)).
113.28(11)° and 131.78(12)°. The third coordination geometry of Cu(II) ions, located on a site with crystallographically imposed twofold symmetry, is represented by the tetrahedrally coordinated Cu2. The Cu2 ion is bound to four different nitrogen atoms (N2, N2a and N8, N8a). The bond length values are 2.008(3) Å for Cu2–N2/2a and 2.027(3) Å for Cu2–N8/8a. The N–Cu2–N angles range from 107.21(11)–116.84(15)°. These values are in very good agreement with those reported in similar literature-known structures (Table 1) featuring Cu(II) triazolate coordination units.13,23

The analysis of the data from single crystal measurement reveals for the metal–organic framework CFA-8 a chemical composition of [Cu2(qpt)]. A schematic representation of the as-synthesized and solvent-filled MOF, which we refer to as large-pore phase CFA-8 network, is presented in Fig. 5. After activation at 250 °C in vacuum all solvent molecules are evaporated out of the pores. The monoclinic network exhibits channel-like pores along the c-axis, which are illustrated as big yellow rods.

The maximum pore opening diameter of the channels calculated with the Poreblazer software24 is around 7.8 Å (Fig. 6, left). The plot on the right side in Fig. 6 along the b-axis shows clearly that there are no other kinds of pores or channels running in different crystallographic directions in the 3D framework. The cyan dashed field illustrates the copper–nitrogen-strands which also run along the c-direction. The channels in the as-synthesized large-pore-phase of CFA-8 are occupied by disordered solvent molecules (DMA), the positions of which were impossible to resolve and refine from the electron density distribution. According to the crystallographic data there is an electron count of 814 per unit cell (rest of the electron density), which correspond to approximately 17 N,N-dimethylacetamide molecules in the unit cell of CFA-8. Calculations with the SQUEEZE program25 reveals that the initial solvent accessible void volume is 3358.4 Å³, which is about 47% of the unit cell volume (7141.0 Å³).

As mentioned before, CFA-8 exhibits permanent porosity. In order to remove all solvent molecules, which are localized in the channels, the as-synthesized sample was activated at 250 °C for at least twelve hours in high-vacuum. Powder XRD measurements show that the unit cell of CFA-8 is changed upon this activation. In the following the heated and solvent-free CFA-8 network is referred to as activated phase of CFA-8. The most obvious difference is the disappearance of the most intensive peak of the XRD pattern around 4.9° in 2θ and the formation of a new peak at 5.9° in 2θ (Fig. 7). The framework remains crystalline after activation but has a smaller unit cell. This process is fully reversible. Thus, treatment of a completely activated sample of CFA-8 with fresh DMA for several hours results in the regeneration of the original as-synthesized phase (= large-pore phase of CFA-8), as shown by XRPD measurements (Fig. 7). Such breathing effects have been repeatedly reported for different MOF structures.26-29

Unfortunately, many attempts to collect a single crystal diffraction pattern of a completely dry single-crystal of CFA-8 were

Table 1 Selected crystallographic and geometric parameters for the 1D Cu-coordination polymer [Cu(bta)] and the 3D CFA-8 [Cu2(qpt)]

<table>
<thead>
<tr>
<th>Coordination polymer, [Cu(bta)]</th>
<th>CFA-8, [Cu2(qpt)]</th>
</tr>
</thead>
<tbody>
<tr>
<td>Wyckoff position</td>
<td>Site symmetry</td>
</tr>
<tr>
<td>Cu1 1f</td>
<td>1</td>
</tr>
<tr>
<td>Cu2 4e</td>
<td>2</td>
</tr>
<tr>
<td>Cu3 4e</td>
<td>2</td>
</tr>
</tbody>
</table>

Selected geometric parameters

| Cu1–N3 | 1.942(3) | Cu1–N1 | 1.931(3) |
| Cu1–N1 | 1.949(3) | Cu1–N3ii | 1.935(3) |
| Cu1–N5 | 2.003(3) | Cu1–N9 | 1.995(3) |
| Cu2–N2 | 2.032(3) | Cu2–N2 | 2.008(3) |
| Cu2–N2i | 2.032(3) | Cu2–N8 | 2.027(3) |
| Cu2–N6 | 2.032(3) | Cu2–N2iv | 2.008(3) |
| Cu2–N6i | 2.032(3) | Cu2–N8iv | 2.027(3) |
| Cu3–N4 | 1.865(3) | Cu3–N10 | 1.874(3) |
| Cu3–N4ii | 1.865(3) | Cu3–N10a | 1.874(3) |

Symmetry codes: (i) x, y, z – 1/2; (ii) –x, y, –z – 1/2; (iii) –x, y, –z – 1/2; (iv) x, –y + 1, z + 1/2; (v) x + 1/2, y + 1/2, z; (vi) –x, y, z; (vii) –x + 1, y, –z + 1/2; (viii) –x + 3, y + 1/2, –z + 1/2; (viii) –x + 1, y, –z + 3/2.
cell volume amounts to 13856 Å³. The unit cell parameters of the as-synthesized CFA-8 (large-pore phase) arises from slight variations of the unit cell parameters. The phase (the same crystal system and space group as for the large-pore phase) of the data of the narrow-pore phase of CFA-8 leads to the treatment of the channels of the porous framework structure, will from now on be referred to as narrow-pore phase, containing trimethyl phosphite guest molecules inside the unit cell parameters. The unit cell parameters of the as-synthesized CFA-8 (large-pore phase) single crystal are $a = 27.49 \text{ Å}$, $b = 23.66 \text{ Å}$, $c = 11.00 \text{ Å}$, $\alpha = 93.75^\circ$ with the cell volume of 7141 Å³. In contrast to this solution, the second crystal (narrow-pore phase) shows lattice parameters of $a = 32.04 \text{ Å}$, $b = 16.79 \text{ Å}$, $c = 10.77 \text{ Å}$, $\alpha, \gamma = 90^\circ$ and $\beta = 92.36^\circ$ with the unit cell volume of 5789 Å³. More details of the narrow-pore phase are presented in the ESI (Table S4; Fig. S9†).

**Thermal analysis and VTXRPD studies**

Structural and thermal stability of CFA-8 were investigated by thermogravimetric analysis (TGA) and variable temperature X-ray powder diffraction (VTXRPD) measurements. In the TGA curve of a pre-dried CFA-8 sample (250 °C/vacuum/10 h) no weight loss is detectable until a temperature of 300 °C is reached (Fig. 8). The framework collapses under dry air (red curve) at around 300 °C, because an exothermal reaction takes place and after decomposition of the organic ligand only copper(II) oxide is left. Under inert nitrogen atmosphere CFA-8 is stable up to 400 °C (blue curve). Above 400 °C CFA-8 is not stable and the decomposition of the framework starts. These results from TG analysis are in very good agreement with VTXRPD measurements (Fig. 9), which were performed in a quartz capillary under inert atmosphere. According to the latter, the CFA-8 network is stable up to 350 °C. In the diffractogram recorded at 400 °C, a new phase with peaks at 46° and 51° in 2θ could be detected, which corresponds to elemental copper (PDF no. 4-836). No peaks of the CFA-8 phase can be seen in the XRPD pattern at a temperature greater than 450 °C.

**Gas sorption measurements**

The adsorption/desorption isotherms for CFA-8 measured at 87.3 K for Ar and at 194.7 K for CO₂ show in both cases a strong hysteresis, which indicates a flexible character of the MOF structure (Fig. 10 and 11). It should be noted, that Ar isotherms show several steps at $p/p_0 > 0.05$ which probably corresponds to the structural transformation in crystals with

![Fig. 8 Thermogravimetric analysis of pre-dried CFA-8 under nitrogen (blue) and synthetic air (red).](image-url)
different sizes. The CO₂ adsorption isotherm shows, in contrast, only one step at p/p₀ > 0.05 corresponding to the structural transformation. This behavior is confirmed by in situ XRPD measurements under CO₂ atmosphere which show structural changes at p/p₀ > 0.1 as well (Fig. S13†). Reversible structural changes of CFA-8 are also shown by in situ XRPD measurements under 80% humidity (Fig. S14†). Since argon sorption isotherms seem to be not equilibrated and dependent on the crystal size at p/p₀ > 0.05 (probably due to the low measurement temperature), the values of the BET surface areas and pore volumes, obtained from CO₂ sorption isotherms, seem to be more reliable. The desorption branch of the CO₂ isotherm in the relative pressure range 0.06–0.1 reveals a BET surface area of 446 m² g⁻¹ for the activated and totally dried CFA-8 phase. The adsorption branch of the Ar isotherm in the p/p₀ range 0.007–0.05 reveals a BET surface area of 518 m² g⁻¹ for the same phase. The values calculated with the Poreblazer software²⁴ (510 m² g⁻¹ for argon with a probe diameter of 3.4 Å and 437 m² g⁻¹ for CO₂ with a probe diameter of 3.8 Å) fit very well. The desorption branch of the CO₂ isotherm in the p/p₀ range 0.06–0.08 reveals a value of 520 m² g⁻¹ for the BET surface area of the large-pore phase. The calculated value from the Poreblazer software for the large-pore phase of CFA-8 is 556 m² g⁻¹ for CO₂. The pore volume of the activated phase determined from the adsorption branch of the CO₂ isotherm at p/p₀ = 0.6 is 0.15 cm³ g⁻¹, the volume of the large-pore phase determined from the adsorption branch of the CO₂ isotherm at p/p₀ = 0.95 is 0.20 cm³ g⁻¹.

Analysis of the argon adsorption isotherm at p/p₀ < 0.05 using the non-local density functional theory (NLDFT) reveals a quite broad distribution of micropores with two maxima at 5.5 and 9.1 Å (Fig. 12).

The isosteric heat of CO adsorption in CFA-8 determined from the adsorption isotherms in the temperature range between 173 K and 193 K is relatively high at low loading (37 kJ mol⁻¹) and decreases to 18 kJ mol⁻¹ at 0.9 mmol g⁻¹ loading (Fig. 13). At higher loading, the isosteric heat of CO adsorption reaches a nearly constant value of about 15–16 kJ mol⁻¹, which corresponds to a typical physisorption value. The value for the isosteric heat of H₂ adsorption, in contrast, remains constant at 6 kJ mol⁻¹ over the whole measured range. An enhanced heat of carbon monoxide adsorption at loading below 0.9 mmol g⁻¹ (16–37 kJ mol⁻¹) indicates a weak chemisorption and coordination of CO molecules to Cu(1) sites. The binding of CO molecules in this case is much weaker than reported previously for Cu(i)-MFU-4large, which forms a very stable CO-complex (stable till 200 °C).³³ Such
weak chemisorption could be advantageous for applications such as gas separation or sensing, because CO binds reversibly at low temperatures and can be easily removed.

As mentioned before, CFA-8 exhibits three different coordinated Cu ions. The value of 0.9 mmol g$^{-1}$ corresponds to one copper(I) atom per formula unit.

Tetrahedrally coordinated Cu$_2$ ions in the CFA-8 framework cannot interact with CO molecules because of their full coordination environment. Apart from binding stoichiometry the synchrotron powder diffraction measurements of CFA-8 with adsorbed CO molecules indicate formation of a Cu I–CO complex between Cu$_1$ (previously trigonally coordinated) and CO (Fig. 14).

The structure of CFA-8 with adsorbed CO molecules was solved from synchrotron powder diffraction data at 100 K. Under CO gas adsorption, the crystal structure of CFA-8 has monoclinic space group C$2/c$ with lattice parameters of $a = 30.866(4)$ Å, $b = 18.881(1)$ Å, $c = 10.851(2)$ Å, $\alpha = 90^\circ$, $\beta = 93.29(2)^\circ$ and a unit cell volume of 6313 Å$^3$, thus having lattice parameters which are between the cell parameters for large-pore and narrow-pore phase of CFA-8. One CO molecule was found to coordinate the Cu$_1$ atom, and additional four molecules were physisorbed in the center of the pores of CFA-8. The originally planar trigonally coordinated Cu$^+$_I atom (Cu$_1$) shows a distorted tetrahedral coordination upon coordination of a single CO molecule. During structure refinement from powder diffraction data, the linear CO molecule was modeled by the single sphere with a position corresponding to the center of gravity of CO molecule. The comparison of structure solution from powder diffraction data with a DFT-calculated model (full lattice relaxation with CASTEP V8.0 PBE-DFT, plane-wave norm-conserving pseudo-potentials, $E$_{cut-off} = 880 eV), shows good agreement between experimentally determined and theoretically calculated positions of CO molecules binding to the corresponding Cu$^+$_I atoms.

The adsorption of CO in CFA-8 was further studied by diffuse reflectance Fourier-transform IR spectroscopy (DRIFT). The recorded IR spectra are shown in Fig. 15. The pre-dried and activated CFA-8 was cooled down to 173 K and the IR spectrum was recorded under nitrogen atmosphere (black curve Fig. 15). After this measurement the N$_2$ atmosphere in the reaction chamber was replaced by CO (5% CO in Ar) for one minute and changed again to N$_2$ atmosphere. The corresponding FT-IR spectrum pictured in light green shows a strong adsorption band at 2088 cm$^{-1}$. The temperature was gradually increased resulting in a decrease of the recorded adsorption band. The additional bands arising at 2173 cm$^{-1}$ and 2117 cm$^{-1}$ belong to free CO molecules in the gas phase. The recorded band at 2088 cm$^{-1}$ corresponds to the stretch mode of the CO molecule coordinatively bound at the Cu$^+$_I site. The vibrational frequency is shifted to lower wavenumber as compared to those of the free gas molecule and stays in good agreement with similar literature-known systems.$^{33,35-37}$

Preliminary examination for catalytic activity

Since the open coordination sits at the coordinatively unsaturated Cu$I$/Cu$_3$ center in the CFA-8 framework could be very interesting for catalytic applications, a simple oxidation test reaction in liquid phase was performed (Scheme 3). The oxidation of ethylbenzene was previously employed for similar purposes in our reports on 3d transition metal-containing metal–organic frameworks MFU-1 and MFU-4$^{15,38}$.

Coordinatively unsaturated Cu$^+$_I ions in a porous framework are promising active sites for catalytic applications. The test
reaction reported here involves NHPI (N-hydroxyphthalimide) as a co-catalyst and represents an effective strategy to activate oxygen from air under mild conditions in the presence of transition metal ions. The so-formed PINO (phthalimide N-oxyl) radical acts as initiator for the radical oxidation of C–H bonds.15 The mechanism of this type of reaction has been described by Ishii et al. in detail.39–43 An induction period at the beginning of the reaction indicates the slow formation of PINO radicals needed as an initiator. The reaction was performed in acetonitrile under oxygen abstracted from the atmosphere. The only detected product of the oxidation reaction of ethylbenzene catalyzed by CFA-8 is acetophenone. The conversion of ethylbenzene reached after 70 h reaction time is around 60% and is nearly equal to the yield of acetophenone determined by GC/MS analysis (Fig. 16). No 1-phenylethanol as a by-product was detected. Very similar values for the conversion of ethylbenzene and the total yield of acetophenone indicate that no over-oxidation and formation of benzoic acid takes place under these mild reaction conditions. The remaining crystallinity of the recovered catalyst was checked after each run (see Fig. S7†). Also the performance of the recycled copper MOF was proven. At least for three runs the activity for the oxidation reaction and also the crystallinity of the CFA-8 network remains the same. The value for BET surface of the catalyst recovered after complete catalytic reaction is about 503 m² g⁻¹ – close to the original value of 518 m² g⁻¹. To confirm the heterogeneous nature of the catalytic test reaction, we performed a hot-filtration experiment, since it was known that even small amounts of leached metal ions in the solution could be a single source of the observed catalytic activity.38 After seven hours, the CFA-8 particles were removed from the reaction mixture by isothermal filtration. The recorded process for the filtered reaction mixture showed no significant conversion after removal of the CFA-8 catalyst.

Conclusion

In conclusion, we have synthesized and structurally characterized a novel metal–organic framework CFA-8 constructed from tppt²⁻ ligands and Cu(i) ions as metal nodes. The Coordination Framework Augsburg University-8 is to the best of our knowledge the first example for a copper(i) MOF incorporating a bistriazolate ligand. A special feature of the here reported structure is its flexible character. The breathing effect, which is common for several MOF structures, is completely reversible as shown by the recorded powder XRD patterns.

Owing to the flexible nature of the framework its gas sorption properties are quite complex. Both adsorption/desorption isotherms (Ar and CO₂) exhibit a strong hysteresis, which is an indication for the flexible character of the network. The calculated heat of adsorption for hydrogen is constant at 6 kJ mol⁻¹ and typical for physisorbed hydrogen. In contrast the value for reversible chemisorbed hydrogen in Cu¹-MFU-4large is 37 kJ mol⁻¹.13 A special focus for the CFA-8 is placed on the sorption properties towards carbon monoxide. From low loading on up to 0.9 mmol g⁻¹ the loading-dependent heat of adsorption values range from 37–16 kJ mol⁻¹. From 0.9 mmol g⁻¹ to higher loading the value is nearly constant at 15–16 kJ mol⁻¹. The weak chemisorption of CO molecules at the unsaturated trigonally coordinated Cu(i) ions is very interesting for applications such as gas separation or sensing and was further confirmed by synchrotron powder diffraction and IR measurements which revealed a characteristic absorption band at 2088 cm⁻¹ corresponding to the Cu¹–CO complex.

A last feature of this novel MOF is the catalytic activity of the framework. A first preliminary examination for the catalytic activity was performed with a simple liquid phase oxidation reaction. The oxidation of ethylbenzene under mild conditions catalyzed by CFA-8 in the presence of NHPI as a co-catalyst leads to acetophenone as a single product. No byproduct, such as 1-phenylethanol or benzoic acid, is detectable. The relatively high conversion of 60% was achieved in this reaction. Further investigations on the sorption properties and the catalytic activity are the research focus in the next time.

Experimental

Materials and methods

All starting materials were of reagent grade and used as received from the commercial supplier. Thermogravimetric analysis (TGA) was performed with a TGA Q500 analyzer in the
temperature range of 25–800 °C in flowing nitrogen and synthetic air at a heating rate of 10 K min\(^{-1}\). Ar, CO, CO\(_2\) and H\(_2\) sorption isotherms were measured with a BELSORP-max instrument combined with a BELCryo system. Adsorbed gas amounts are given in cm\(^3\) g\(^{-1}\) [STP], where STP = 101.3 kPa and 273.15 K. Prior to measurements, the sample was heated at 250 °C for 20 h in high vacuum in order to remove occluded solvent molecules. Fourier transform infrared (FTIR) spectra were recorded with ATR in the range 400–4000 cm\(^{-1}\) on a Bruker Equinox 55 FT-IR spectrometer. Energy-dispersive X-ray spectroscopy (EDX) was performed with a Philips XL 30 FEG scanning electron microscope equipped with an EDAX SiLi detector. Elemental analyses were measured with a Vario EL III instrument from Elementar Analysensysteme GmbH. GC/MS measurements were performed on a Hewlett Packard GC System 6890 Series equipped with Rtx-5MS column and mass selective detector MSD 5973. Powder X-ray diffraction data were collected in the \(2\theta\) range with 0.02° steps, with a time of 3 s per step, by using a Seifert XRD 3003 TT diffractometer equipped with Meteor 1D detector. The diffuse reflectance Fourier-transform IR spectra (DRIFT) were collected between 3500–400 cm\(^{-1}\) on an Equinox 55 FT-IR spectrometer equipped with a Praying Mantis diffuse reflectance accessory and an environmental chamber (Harrick Scientific Products) and referenced to KBr. The H\(_2\)-tqpt ligand was synthesized in three steps from commercially available 5,5-dimethylcyclohexane-1,3-dione. The following methylation and oxidation reactions steps from commercially available 5,5-dimethylcyclohexane-1,2,4,5-tetraone (1.00 g, 5.10 mmol) and 1H-benzotriazole-5,6-diamine (1.60 g, 10.70 mmol) in 200 mL of methylacetamide (DMA) and the solution was placed in a glass tube (10 mL). The tube was sealed and heated at 140 °C for 2 days and then cooled to room temperature. The orange precipitate was filtered, washed with DMA (3 × 3 mL) and dried at 250 °C under vacuum.

Yield: 7.99 mg (73%). IR (\(\nu\) (cm\(^{-1}\))): 2968, 1456, 1371, 1272, 1211, 1146, 1105, 850, 806, 613, 426. Elemental analysis calcd (%): C 48.26, H 2.95, N 25.58; found: C 48.29, H 2.99, N 25.70.

N.B.: all trials to synthesize CFA-8 from Cu(i) salt (Cu(i)Cl, Cu(i)Br or Cu(i)Ac) under similar reaction conditions have failed in our hands so far. Similarly, replacing DMA by other organic solvents like \(N, N\)-dimethylformamide (DMF), \(N, N\)-diethyl-formamide (DEF) or \(N\)-methylformamide (NMF) did not yield phase pure CFA-8.

**Synthesis of CFA-8 (narrow-pore phase)**

An as-synthesized sample of CFA-8 was activated for 20 hours in vacuum at 250 °C. Subsequently the powder sample (15 mg) was suspended in a methanolic solution of trimethyl phosphite (20 mL, 0.3 mmol mL\(^{-1}\)). The mixture was stirred for 30 min, filtered and washed with fresh MeOH. The still orange powder was dried for 4 hours at 60 °C per vacuum.

**Oxidation of ethylbenzene with CFA-8**

For preliminary examination for catalytic activity of the Cu(i)-MOF, an easy oxidation reaction was performed with CFA-8 as a catalyst and \(N\)-hydroxyphthalimide (NHPI) as a co-catalyst. The XRPD data under CO\(_2\) pressure and relative humidity were collected with an Empyrean (PANalytical) Diffractometer equipped with Bragg-Brentano\(^{\text{TM}}\) mirror, PIXcel\(^{\text{TM}}\) 2 x 2 detector and Cryo & Humidity Chamber CHC plus\(^{\text{TM}}\) (Anton Paar).

To a solution of ethylbenzene (1.06 g, 10 mmol), NHPI (163 mg, 1.0 mmol) and \(n\)-decane (0.71 g, 5 mmol, used as a standard) in acetonitrile/water mixture (98:2, 40 mL) was added an activated sample of CFA-8 (27 mg, 0.24 mmol). The suspension was stirred at 40 °C. To analyze the conversion of the reaction 0.1 mL of the mixture was filtered through neutral Al\(_2\)O\(_3\) which was then eluted with Et\(_2\)O/CH\(_2\)Cl\(_2\)/MeOH (10 : 10 : 1) and GC/MS measurements were repeated three times for each sample. Ethylbenzene and acetophenone were used as standards for the determination of the conversion and yield. To check the performance and the recyclability of the CFA-8 MOF the catalysis experiment was repeated three times with the same CFA-8 sample recovered after the reaction. After each experiment the XRPD pattern of the framework was examined.

**Single-crystal X-ray diffraction**

For the single-crystal X-ray measurements, the crystals of [Cu\(_2\)(tqpt)]\(\cdot\)4.25DMA (large-pore phase) were taken from the mother liquor, while the crystals of [Cu\(_2\)(tqpt)]\(\cdot\)0.94(CH\(_3\)O\(_3\))\(_2\) (narrow-pore phase) were dried in vacuum (4 h/60 °C) prior to the measurements. The refined solution of the crystals with the composition [Cu\(_2\)(tqpt)]\(\cdot\)0.94(CH\(_3\)O\(_3\))\(_2\) describes the crystal system for the narrow-pore phase of CFA-8. These crystals were obtained by an experiment in order to check, if the phosphorus atom of trimethyl phosphate coordinates to the Cu1 atom as observed for CO gas molecules. In this case, the interactions between the P donor atom and the Cu1 atom are not strong enough, but the single crystal structure of the narrow-pore phase of CFA-8 could be solved. The crystals were mounted on a MiTeGen MicroMount and tested on a Bruker D8 Venture diffractometer. X-ray diffraction data for the single crystal structure analysis were collected on a Bruker D8 Venture dif
Table 2  Crystal and experimental data for the large-pore phase of CFA-8

<table>
<thead>
<tr>
<th></th>
<th>CFA-8-4.25 DMA, large-pore phase</th>
</tr>
</thead>
<tbody>
<tr>
<td>Empirical formula</td>
<td>C_{61}H_{70.25}Cu_{4}N_{24.25}O_{4.25}</td>
</tr>
<tr>
<td>Formula weight</td>
<td>1465.32</td>
</tr>
<tr>
<td>Temperature (K)</td>
<td>100(2)</td>
</tr>
<tr>
<td>Wavelength (Å)</td>
<td>0.71073</td>
</tr>
<tr>
<td>Crystal system</td>
<td>Monoclinic</td>
</tr>
<tr>
<td>Space group</td>
<td>C2/c (no. 15)</td>
</tr>
<tr>
<td>Unit cell dimensions (Å)</td>
<td>a = 27.4916(12) Å, b = 23.6639(11) Å, c = 11.0002(5) Å, β = 93.7500(14)°</td>
</tr>
<tr>
<td>Volume</td>
<td>7141.0(6) Å³</td>
</tr>
<tr>
<td>Z</td>
<td>4</td>
</tr>
<tr>
<td>Density (calculated) (g cm⁻³)</td>
<td>1.363</td>
</tr>
<tr>
<td>Absorption (mm⁻¹)</td>
<td>1.237</td>
</tr>
<tr>
<td>f(000)</td>
<td>3024</td>
</tr>
<tr>
<td>Crystal size (mm³)</td>
<td>0.117 × 0.04 × 0.03</td>
</tr>
<tr>
<td>Theta range for data collection (°)</td>
<td>2.27 to 25.09</td>
</tr>
<tr>
<td>Refls. collected</td>
<td>38185</td>
</tr>
<tr>
<td>Refls. unique</td>
<td>6325 [R (int) = 0.0867]</td>
</tr>
<tr>
<td>Data/restraints/parameters</td>
<td>6325/0/312</td>
</tr>
<tr>
<td>GooF</td>
<td>1.047</td>
</tr>
<tr>
<td>R₁ (I &gt; 2σ(I))</td>
<td>0.0460</td>
</tr>
<tr>
<td>wR₂ (all data)</td>
<td>0.1171</td>
</tr>
<tr>
<td>Largest diff. peak and hole (e Å⁻³)</td>
<td>0.734 and −0.829</td>
</tr>
</tbody>
</table>

VT-X-ray powder diffraction

VT-XRPD measurements were performed with a Bruker D8 Advance θ–2θ diffractometer in transmittance Bragg–Brentano geometry employing a Göbel mirror, Cu-radiation and equipped with a LynXeye 1-D detector. The sample was ground in an agate mortar and loaded into quartz capillaries (Hilgenberg) with 0.5 mm diameter and 0.01 mm wall thickness. The patterns were recorded in a temperature range of 30 to 550 °C, in the 3–60° 2θ range, with a step time of 1 s and a step width of 0.02° 2θ. Temperature program between measurements: 0.5 °C s⁻¹ heating rate, followed by 10 min isothermal steps required for recording diffraction data sets.

Variable-temperature synchrotron X-ray powder diffraction measurements of evacuated CFA-8 were performed at temperatures 500 K, 270 K, 170 K, 130 K, 100 K at beamline I12 at Diamond (UK). Synchrotron X-ray powder diffraction measurements of CO adsorption in CFA-8 were performed at 100 K, 100 kPa gas pressure. Sample was evacuated before gas loading, cooled down to 100 K, and exposed to CO gas.

Indexing and structure solution by simulated annealing of the crystal structure of CFA-8 with adsorbed CO molecules was performed using TOPAS. Subsequent Rietveld refinement confirmed the crystal structure.

Acknowledgements

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Notes and references

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