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# Synthesis of hierarchically structured materials: microporous diatoms and nanoporous hydroxyaluminosilicate

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**Abstract** Diatomite is well known to remove particulate matter for a broad range of applications due to diatom skeleton structures. Hydrothermally treated diatomite has been demonstrated to be a feasible pathway to produce advanced materials for the removal of aqueous heavy metal ions. However, single-synthesis routes to obtain altered diatomite for simultaneous removal of dissolved and particulate contaminants are still missing. Therefore, hydrothermal experiments at pH 12.8 ([KOH] = 0.1 M), and  $T = 100\text{ °C}$  and different added Al concentrations ( $0\text{ mM} \leq [\text{Al}] \leq 11\text{ mM}$ ) were conducted. Results reveal that in comparison with previous experiments (Höllén et al. in *J Mater Sci Eng B* 2(10):523–533, 2012) using less alkaline solutions slower dissolution of diatomite is triggering the formation of nanosized spherical particles (NSP) in the presence of diatom skeleton relicts. NSP ranged between 10 and 50 nm in size and were located within the

micropores of the diatoms (100–200 nm). TEM-SAED analyses of NSP indicated nanocrystalline phases with distinct lattice planes of 0.45 and 0.26 nm, which could not be attributed to common aluminosilicates. In accordance with the small crystal size, no X-ray diffraction peaks occurred in the bulk reaction product using powder XRD analyses. Metal ion removal experiments at ambient temperature indicated its high removal capacities for aqueous  $\text{Cu}^{2+}$ ,  $\text{Pb}^{2+}$  and  $\text{Zn}^{2+}$  ranging from 10 to 99 % (e.g., ion sorption capacities of about 0.02–0.2 mval/g). The synthesized product was also successfully tested for its metal ion removal behaviour using a mine drainage solution containing  $\text{Cd}^{2+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Zn}^{2+}$  and  $\text{Sr}^{2+}$ . The developed synthesis path yields a hierarchically structured material based on combined micropores (about 100 nm) from the diatoms and NSP with nanopores (1–5 nm) which is promising for large-scale production to simultaneously remove dissolved and particulate contaminants from aqueous solutions.

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## Introduction

Diatomite is composed of siliceous algae [1] which are widely used for separation of  $\mu\text{m}$ -sized suspended particles [2] and as substrate for the synthesis of siliceous reaction products, like mordenite [3], analcime, cancrinite and hydroxysodalite [4], for ion removal from aqueous solution. Nevertheless, synthesis routes to obtain tailored hierarchically structured altered diatomite for a

simultaneous removal of dissolved and particulate contaminants are rare and still challenging with respect to feasibility and cost efficiency.

For instance, a hierarchically structured diatomite/MFI-type zeolite composite was produced by using a NaOH solution for etching the pristine diatomite and adding nanocrystalline silicalite-1 seeds, which was successfully tested for benzene adsorption [5]. The synthesis of these types of composite materials fits well to similar approaches using various types of templates [6]. An alternative is the synthesis of hollow nanozeolites using organic templates (polydiallyldimethylammonium chloride or polyurea/polyformaldehyde) which is removed by calcination [7, 8]. The use of mesoporous silica spheres as template [9] allows abstaining from calcination. Thus, hierarchically structures were successfully synthesized in former studies using synthetic materials like polystyrene or carbon black [10], but limited experiences exist with respect to using naturally occurring materials which offer micro-sized pre-structures for the above approach. A promising material is in particular diatomite from natural deposits which has been used for a long time for filtration (water treatment, beer production, etc.) and is available at low costs and large volumes. In previous studies, hydrothermally treated diatomite yielded mostly zeolites, where experiments without and with seeds (e.g., silicalite nanoparticles) were conducted ([11, 12], respectively). In our former experimental approach [11], we conducted experiments using strong alkaline solutions with or without additional Al for hydrothermal alteration of diatomite which yielded in the formation of zeolite throughout the formation of an intermediate phase. The latter revealed an extraordinarily high metal ion removal capacity. However, synthesis of a hierarchically structured material containing both micro- and nanopores could not be realized due to fast diatomite dissolution and thus losing microporous structures.

The aim of the present study is to obtain both micro- and nanopores within a single synthesis and reaction product and to test its suitability for ion removal. For this purpose, an experimental approach for hydrothermal alteration of diatomite using moderate alkaline solutions and a temperature of 100 °C without using any seed material or template was chosen. The product is characterized with respect to its chemical, mineralogical and micro-/nanostructural characteristics and tested for its ion removal behaviour.

## Experimental

### Analytics

The chemical composition of aqueous solutions was analysed by inductively coupled plasma mass spectroscopy

(ICP-OES; Perkin Elmer Optima 4300) with an accuracy of  $\pm 5\%$  and a precision of  $\pm 2\%$  for Si, Al, K, Ca, Mg, Cd, Cu, Pb and Zn. ICP-MS (Agilent 7500 cx) was used for quantification of Sr and Cs with an accuracy of about  $\pm 10\%$  and a precision of about  $\pm 5\%$ . The pH was analysed with a precision of  $\pm 0.03$  using WTW Inolab 740 with the pH electrode BL 26 (Schott) using a three-point calibration (pH 4, 7 and 10).

Solids were mineralogically characterized by X-ray diffraction pattern (XRD, PANalytical X'Pert Pro; Co K $\alpha$ , 40 mA, 45 kV, sample rotation) between  $4^\circ$  and  $110^\circ$  ( $2\theta$ ). Secondary electron microscopy images of solids were obtained by scanning electron microscopy (Zeiss DSM 982 Gemini) at an accelerating voltage of 5 kV and working distances between 5 and 10 mm. Therefore, samples were coated with an Au–Pd alloy using a Scancoat Six sputter apparatus (Edwards Hochvakuum GmbH). Transmission electron microscopy (TEM, FEI Tecnai 12 for TEM-SAED, Philipps CM20 for TEM-EDX) was used for chemical and structural characterization at a nanoscale. The specific surface area and related pore size distribution calculations are based on gas adsorption analyses. The specific surface area ( $S$ ;  $\pm 5\%$ ) was measured by the BET method [13] using a FlowCash II 2300 equipment and a nitrogen–argon mixture. For selected samples, the mesopore radii distribution curve was determined by BJH method [14] using a Micromeritics TriStar II 3020 V1.03. With this instrument also the micropore volume could be derived from the  $t$ -plot approach [15].

Hydrogeochemical modelling using PHREEQC with database PhreeqC.dat [16] was used to calculate activities and in the experimental solutions and saturation indices of dissolving (e.g., gibbsite) or possibly precipitating phases. The saturation index is defined as

$$SI = \log \frac{IAP}{K},$$

where IAP stands for the ion activity product and  $K$  for the solubility product.

### Materials

Commercially available unconsolidated and untreated natural diatomite (trade Thiele) from a northern African deposit was used for hydrothermal alteration experiments. Experiments were carried out by using two different solutions: Solution A consists of 0.1 M KOH freshly prepared in a gas-tight Nalgene<sup>®</sup> polycarbonate bottle to prevent CO<sub>2</sub> adsorption (pH = 12.9). Solution B was prepared by suspending 1 g of gibbsite ( $\gamma$ -Al(OH)<sub>3</sub>, Merck p.a.) in 0.5 L of 0.1 M KOH. After a reaction time of 7 days for the dissolution of gibbsite, the reacted solution was separated from the partly remaining gibbsite by

membrane filtration (0.45  $\mu\text{m}$ ). The measured Al concentration of 11 mM represents still undersaturation with respect to gibbsite at the analysed pH of 12.8 ( $SI_{\text{gibbsite}} = -0.33$ ). Due to the increase in the solubility of gibbsite and amorphous aluminium hydroxide (higher soluble compared to gibbsite) with temperature [17], precipitation of  $\text{Al}(\text{OH})_3$  can be ruled out during heating up to 100  $^\circ\text{C}$ .

For cation removal experiments, two different reference solutions were used:

1. A synthetic heavy metal stock solution containing each 0.5 mM of  $\text{Cu}^{2+}$ ,  $\text{Pb}^{2+}$  and  $\text{Zn}^{2+}$  which was prepared by dissolving  $\text{Pb}(\text{NO}_3)_2$ ,  $\text{Zn}(\text{NO}_3)_2 \cdot 4 \text{H}_2\text{O}$  and  $\text{Cu}(\text{NO}_3)_2 \cdot 3 \text{H}_2\text{O}$  (Merck, p.a.) in MilliQ water (pH 5).
2. A mine drainage solution from a carbonate-hosted lead–zinc deposit (Carinthia, Austria). The 0.45  $\mu\text{m}$  filtrated stock solution contains 1.250 mM of Ca, 0.00014 mM of Cd, 0.026 mM of K, 0.665 mM of Mg, 0.153 mM of Na, 0.00048 mM of Pb, 0.0031 mM of Si, 0.00062 mM of Sr and 0.0017 mM of Zn (pH 8).

## Experimental setup

The experimental setup for hydrothermal alteration of diatomite was similar to that in our previous study [11], but differs from it with respect to the molarity of the alkaline solution (0.1 M instead of 1 M KOH) and by focussing on experiments with an alteration temperature of 100  $^\circ\text{C}$  which turned out to be ideal to produce the

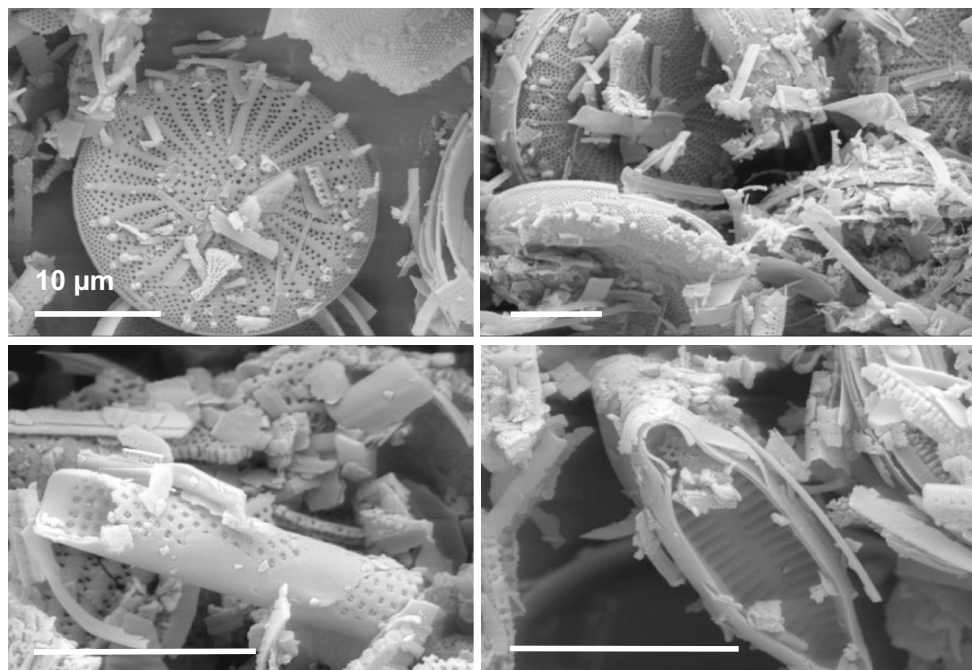
intermediate phase in the former experiments [11]. Briefly, for alteration experiments 0.5 g diatomite reacted with 0.025 L of experimental solutions A or B in a Teflon-coated steel autoclave (0.1 L) at 25 and  $100 \pm 2$   $^\circ\text{C}$ . After distinct reaction times between 0.25 and 64 days, the autoclaves were cooled down to about 30  $^\circ\text{C}$  and the experimental solutions were immediately separated from the obtained solids by membrane filtration (0.45  $\mu\text{m}$ ) for solid and liquid analyses. The solids were rinsed with 0.2 L MilliQ water and dried at 40  $^\circ\text{C}$  before analyses.

The cation removal experiments were carried out by suspending the pristine diatomite material or the reaction products from alteration experiments in the different reference solutions. Therefore, 0.05 g of the reaction product was suspended in 0.01 L of (1) the  $\text{Cu}^{2+}$ ,  $\text{Pb}^{2+}$  and  $\text{Zn}^{2+}$  containing solution and (2) the mine drainage stock solution. The suspension was placed in a horizontal shaking device (E.B. KS-15) with a frequency of  $250 \text{ min}^{-1}$  for up to 3 days at 25  $^\circ\text{C}$ . At distinct time intervals, samples were gathered from the solution for analyses.

## Results

### Pristine diatomite

The pristine diatomite in Fig. 1 is composed of diatoms of the *Centrales* group of mostly disc-shaped form. It consists



**Fig. 1** SEM-SE images of unaltered diatomite showing central and minor amounts of pennate diatoms

of 97 % opaline silica with an L.O.I. of 3.5 wt%. The specific surface area of the diatomite accounts for  $S = 15.5$  and  $20.1 \text{ m}^2 \text{ g}^{-1}$  according to the BET and Langmuir

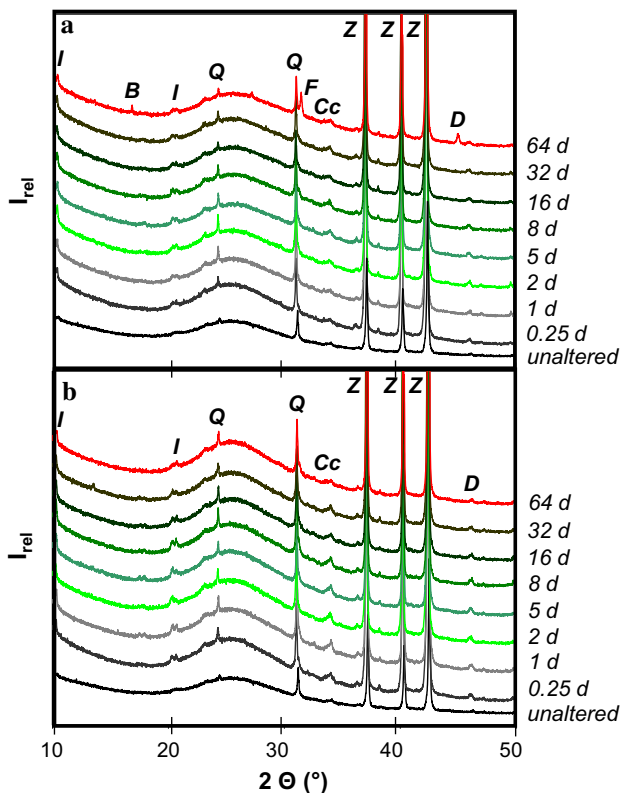
method, respectively. Data derived from  $t$ -plot analyses indicate that  $12.2 \text{ m}^2 \text{ g}^{-1}$  are external surface areas and  $3.3 \text{ m}^2 \text{ g}^{-1}$  micropores. According to BJH adsorption data,  $9.3 \text{ m}^2 \text{ g}^{-1}$  are allotted to a pore diameter range from 1.7 to 300 nm. Values for the average pore diameter range are 13.7 and 20.3 nm for the BET and BJH method, respectively.

### Hydrothermally altered diatomite

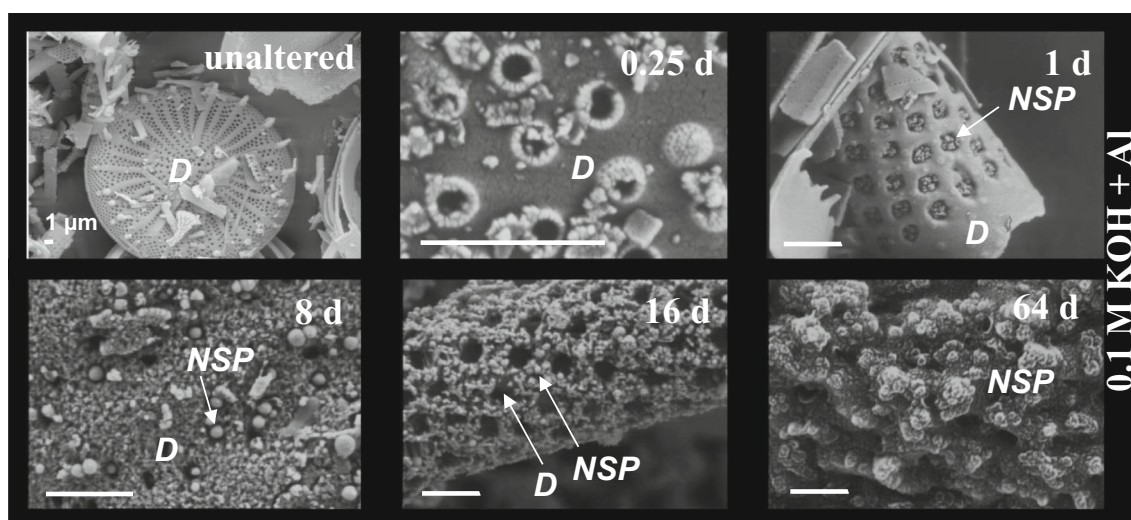
The products from the alteration experiments did not show any X-ray-detectable crystalline phase throughout the whole experimental runs (Fig. 2). However, changing the experimental parameters compared to [11] led to a decrease in the dissolved silica concentration from  $261 \pm 23 \text{ mM}$  using 1 M KOH containing 22 mM Al to  $52 \pm 5 \text{ mM}$  using 0.1 M KOH containing 11 mM Al, each after an initial period of 1 day, which correlates to a less alkaline pH (12.8 for 0.1 M KOH, 13.7 for 1 M KOH). A higher dissolved silica concentration was obtained when using pure 0.1 M KOH which confirms our previous study and is explained by the retarding effect of Al on silica dissolution [18].

However, TEM-SAED patterns of crystalline particles in diatomite altered for either 1 or 16 days show Bragg reflections representing lattice spacings of 0.45 and 0.26 nm. The absence of crystalline features in XRD patterns can be explained by the small size of the crystals and their low percentage in the overall volume of the sample. Further research is needed to identify the crystalline phase which has formed during our experiments.

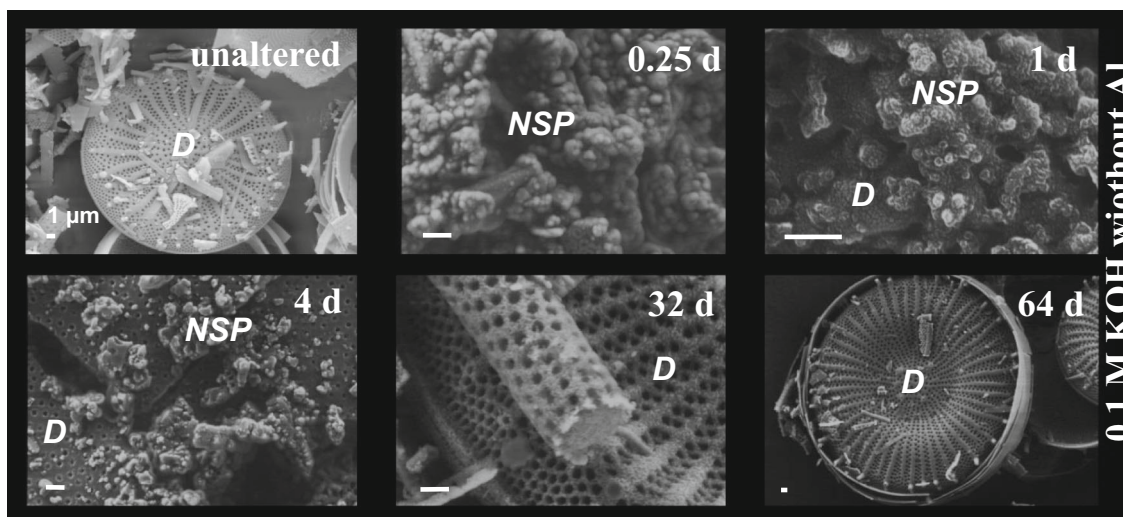
SEM imaging clearly indicates a proceeding dissolution of diatoms (Fig. 3). Additionally, the precipitation of the



**Fig. 2** XRD patterns of alteration products of diatomite in 0.1 M KOH with (a) and without (b) 0.11 mM Al. *I* Illite, *Q* quartz, *Cc* calcite, *Z* zincite (added as internal standard for the quantification of the amorphous content)



**Fig. 3** SEM-SE images of unaltered diatomite and diatomite altered at 100 °C for various reaction times in 0.1 M KOH containing 11 mM of Al. *D* diatom, *NSP* nanosized spherical particles

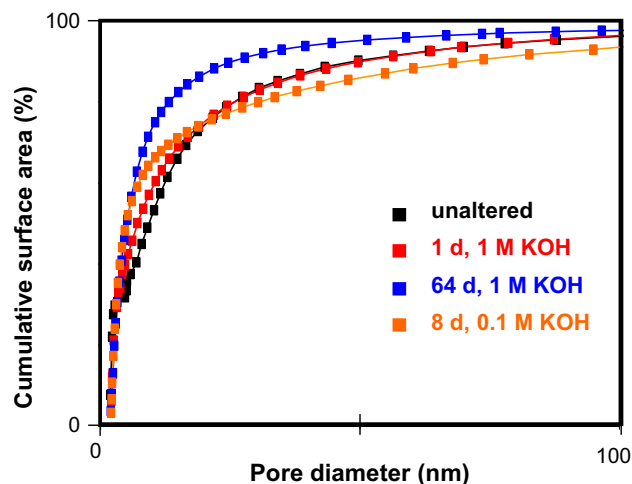


**Fig. 4** SEM-SE images of unaltered diatomite and diatomite altered at 100 °C for various reaction times in 0.1 M KOH without added Al. *D* diatom, *NSP* nanosized spherical particles

spherical-shaped particles with a diameter of 100 nm takes place. These particles were formed directly in the pores and at the surfaces of the diatoms. The formation of these particles occurs in several steps: SEM-SE images indicate already after 0.25 days in experiments with 11 mM Al bulges around the pores of the diatoms. Subsequently, a hierarchically structured material is formed which is composed of these particles in the pores of the diatoms and the diatoms themselves. Later on the spherical particles cover the whole surface of the diatoms, but are not visible within the pores any more. Finally, the surface of the diatoms becomes rougher and remains covered with the particles (Fig. 3).

Using experimental solution A, without dissolved Al, irregularly shaped isometric particles with a diameter of some 100 nm are formed already after 0.25 day (Fig. 4). After 1 day, the material looks similar as after 64 days in the experiments with Al (Fig. 3). This stage might again be described as hierarchically structured material. The particles start to disappear afterwards and are absent after 32 day.

The specific surface area of diatomite increases from  $S = 15.5\text{--}25.4 \text{ m}^2 \text{ g}^{-1}$  during the 8 days of hydrothermal alteration at 100 °C in 0.1 M KOH (experimental solution B). The pore size distribution sum curve (Fig. 5; Table 1) shows that the relative amount of small pores increases with ongoing alteration in 0.1 M KOH, which can be expressed by a decrease in the median pore diameter from 9 to 4 nm within 8 days. On the other hand, the higher relative percentage of larger pores indicates the survival of diatoms. Additional data from a previous study [11] using 1 M KOH are also displayed in Fig. 5 and show a



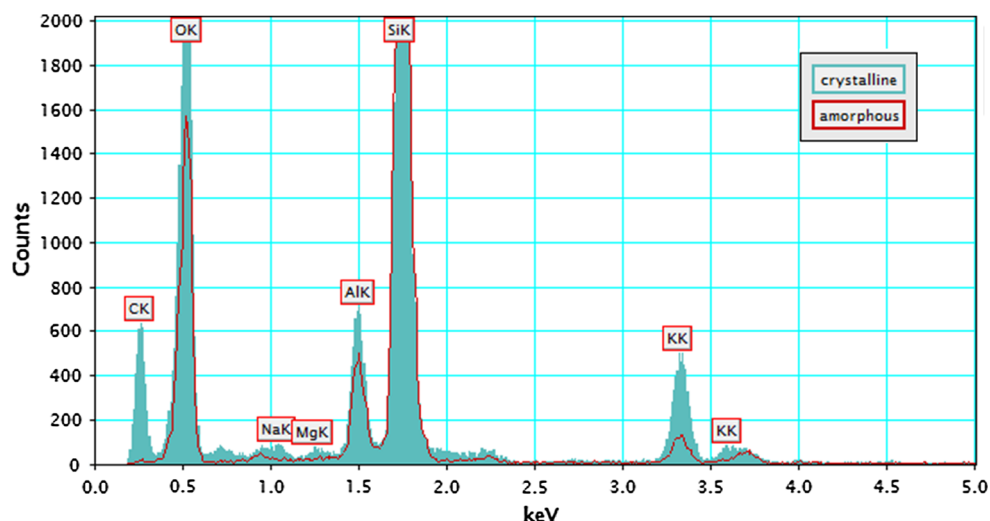
**Fig. 5** Cumulative pore surface area (%) of unaltered diatomite and diatomite altered 8 days in 0.1 M KOH with 11 mM Al (this study) and for 1 and 64 days in 1 M KOH with 22 mM Al at 100 °C [11] as a function of pore diameter analysed by the *t*-plot method by BET

comparable increase in the cumulative pore surface area for pores <10 nm, but a decrease for pores >10 nm.

In Fig. 6, representative TEM-EDX spectra of amorphous and crystalline pore fillings in altered diatomite are shown. TEM analyses of diatomite altered for 1 day in 0.1 M KOH containing Al indicate that the diatoms themselves remain amorphous, but pore fillings are formed which are partially crystalline and enriched in Al compared to the diatoms. Some of these pore fillings contain significant amounts of potassium. Generally, crystalline pore fillings are richer in Al and K than amorphous ones. Some non-spherical particles are randomly oriented, whereas

**Table 1** Incremental pore surface area ( $\text{m}^2/\text{g}$ ) of unaltered diatomite and diatomite altered 8 days in 0.1 M KOH with 11 mM Al (this study) and for 1 and 64 days in 1 M KOH with 22 mM Al at 100 °C [11] as a function of pore diameter analysed by the  $t$ -plot method by BET

Pore range diameter (nm)	Unaltered diatomite ( $\text{m}^2/\text{g}$ )	Diatomite altered 8 days in 0.1 M KOH with 11 mM Al ( $\text{m}^2/\text{g}$ )	Diatomite altered 1 day in 1 M KOH with 22 mM Al ( $\text{m}^2/\text{g}$ )	Diatomite altered 64 days in 1 M KOH with 22 mM Al ( $\text{m}^2/\text{g}$ )
2–10	4.91	14.81	14.72	17.75
10–20	2.10	2.10	4.37	3.22
20–30	0.67	1.13	1.56	1.13
30–40	0.41	0.80	1.11	0.54
40–50	0.26	0.60	0.71	0.31
50–60	0.17	0.51	0.51	0.20
60–70	0.15	0.39	0.42	0.15
70–80	0.10	0.32	0.27	0.12
80–90	0.08	0.24	0.24	0.08
90–100	0.07	0.25	0.19	0.06
>100	0.36	1.47	0.91	0.56

**Fig. 6** Comparison of characteristic TEM-EDX spectra of amorphous and crystalline pore fillings in altered diatomite

others show a preferred orientation. Those particles with a preferred orientation are enriched in Si and Al, whereas the randomly oriented particles are enriched in K and Ca.

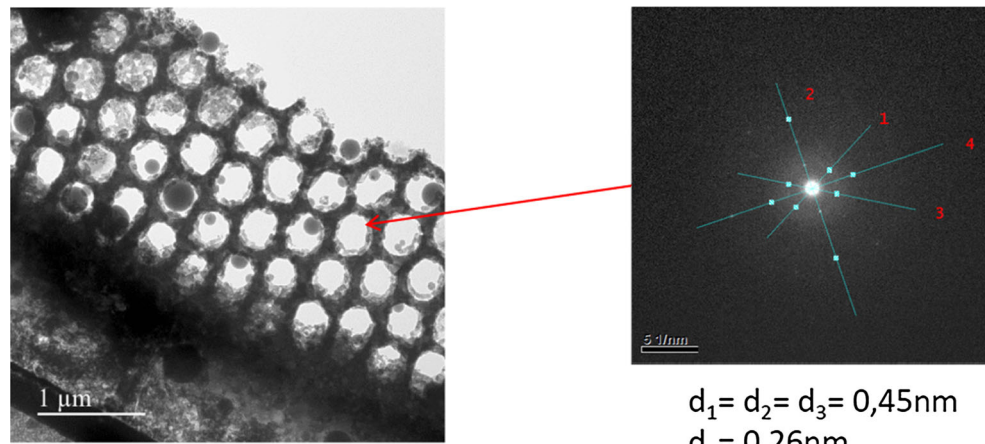
After 16 days in 0.1 M KOH containing 11 mM Al, amorphous spherical particles occur which coexist with irregularly shaped, partly crystalline particles (Fig. 7) known from less altered diatomite (1 day). The spherical particles become enriched with respect to K and Al with increasing size. Crystalline non-spherical particles contain more K and Al than amorphous particles of the same shape. Consequently, increase in structural ordering is associated with an increase in K and Al concentration. Crystalline non-spherical particles are richer in K and poorer in Al compared to spherical, amorphous particles. However, the K content of spherical amorphous particles varies a lot. The Al concentration in crystalline particles is lower than that in large spherical amorphous particles. A comparison

between the chemical composition of pure diatoms and diatoms covered with particles clearly indicates that the particles are enriched in Al, Ca and K.

### Metal removal experiments

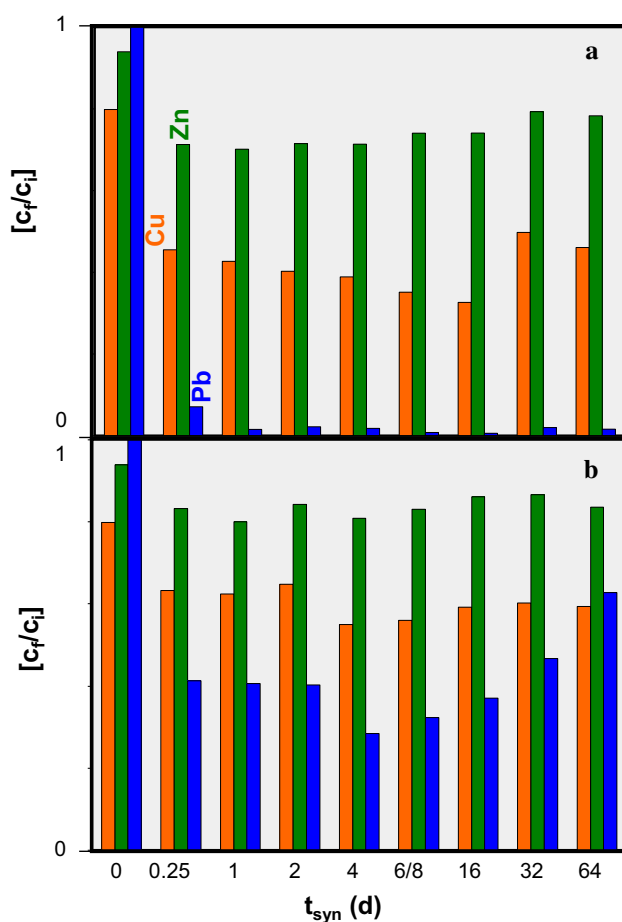
In Fig. 8, the dissolved concentrations of Pb, Cu and Zn after metal removal experiments are presented as fractions of their initial concentration in the synthetic heavy metal solution which accounted for about 0.5 mM. Obviously, unaltered diatomite has only a limited capacity to remove heavy metal ions from the synthetic solution containing Cu, Pb and Zn ions. Pb ion removal is insignificant, and Cu removal accounts for less than 20 %. Alteration of diatomite in 0.1 M KOH creates a capacity to remove Pb ions from the synthetic Cu–Pb–Zn solution. Pb ion removal is more efficient (up to 99 %) when 11 mM Al is present. In the

**Fig. 7** Characteristic TEM image of crystalline nanoparticles in the pores of altered diatoms (left) and evaluation of SAED patterns of the respective particle (right)



$$d_1 = d_2 = d_3 = 0,45\text{nm}$$

$$d_4 = 0,26\text{nm}$$

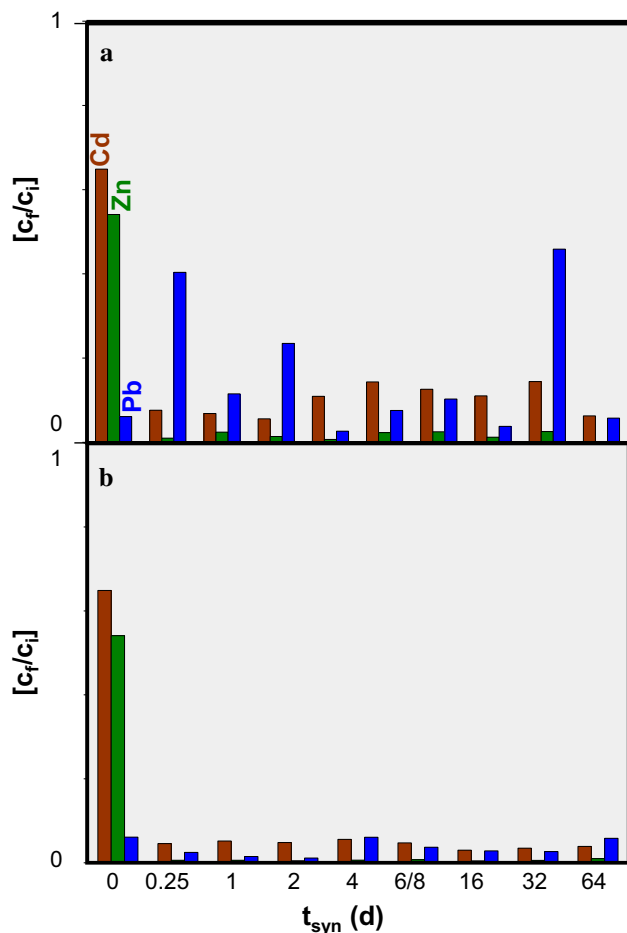


**Fig. 8** Heavy metal ion removal from the synthetic Cu–Pb–Zn solution (each 0.5 mM Cu, Pb and Zn, pH 5) after 3 days of reaction with unaltered diatomite ( $t_{\text{syn}} = 0$  day) and diatomite altered for various durations ( $0.25 \text{ day} \leq t_{\text{syn}} \leq 64$  days) in 0.1 M KOH with or without 11 mM Al (a and b, respectively). Concentrations of dissolved metal ions are given as ratios between the final concentration,  $c_f$ , and the initial concentration,  $c_i$

absence of Al, it only accounts for 40–70 % with a maximum for diatomite altered for intermediate periods (4 days). Similarly, the capacity for removal of Cu ions increases

during alteration of diatomite from 20 % for unaltered diatomite to 40 % during alteration in pure 0.1 M KOH and even to 60 % in the presence of Al. Compared to Cu and Pb, Zn ions are only removed to a very limited degree (10–20 %) from the synthetic Cu–Pb–Zn solution even by altered diatomite. The order of selectivity of heavy metal ion removal from the synthetic solution is therefore  $\text{Pb} > \text{Cu} > \text{Zn}$ , no matter if Al has been present in alteration experiments. Like in our previous study [11], heavy metal removal was associated with release of K from the solid into the solution. Correspondingly, to the lower heavy metal removal also the K release was with resulting dissolved concentrations of  $1.4 \pm 0.1$  and  $0.6 \pm 0.1$  mM for Al-containing and Al-free alteration experiments, respectively, only about one-third of the value obtained previously for alteration productions synthesized using 1 M KOH.

In contrast to the experiments with the synthetic solution containing Cu, Pb and Zn ions, unaltered diatomite removes 94 % of Pb ions and more than 50 % of Zn ions from the mine drainage solution (pH 8) (Fig. 9). Also Cd concentrations could be reduced by 36 % to  $0.009 \mu\text{mol L}^{-1}$ . Alteration products remove heavy metal ions from the mine drainage solution more efficiently than unaltered diatomite. Like for all other alteration products, heavy metal ion removal is connected with K release. The relative amount of heavy metal ions removed by altered diatomite is higher for the mine drainage solution than for the synthetic solution containing Cu, Pb and Zn ions which had much higher initial concentrations. Interestingly, diatomite altered in pure 0.1 M KOH was more capable of removing heavy metals than diatomite altered in Al-containing 0.1 M KOH. Zn ions are removed more efficiently than Pb ions from the mine drainage solution, although Pb ions are removed more efficiently than Zn ions from the synthetic solution containing Cu, Pb and Zn ions. The capacity of Cd ion removal is between that of Pb and Zn ion removal. Generally, the duration of alteration had no impact on the capacity of heavy metal removal. Water



**Fig. 9** Heavy metal removal from a mine drainage solution ( $0.14 \mu\text{mol L}^{-1}$  Cd,  $0.48 \mu\text{mol L}^{-1}$  Pb,  $17 \mu\text{mol L}^{-1}$  Zn, pH 8) after 3 days of reaction with unaltered diatomite ( $t_{syn} = 0$  day) and diatomite altered for various durations ( $0.25 \text{ day} \leq t_{syn} \leq 64$  days) in 0.1 M KOH with or without 11 mM aqueous Al (**a** and **b**, respectively). Concentrations of dissolved metal ions are given as ratios between the final concentration,  $c_f$ , and the initial concentration,  $c_i$

hardness in terms of Ca and Mg concentration was also significantly reduced by altered diatomite, especially when Al had been present in alteration experiments. Diatomite dissolution during metal ion removal experiments as expressed by concentrations of dissolved Si is slightly higher when alteration had taken place in presence of Al.

## Discussion

### Diatomite alteration experiments

In the diatomite alteration experiments, a hierarchically structured material has been formed. A comparison with previous experiments using 1 M KOH [11], where 70–80 wt% of the added diatomite were dissolved, shows

not only slower dissolution kinetics, but also a lower thermodynamic solubility at lower KOH molarity. This is in agreement with a similar study on hydrothermal dissolution of diatomite due to replacement of Si by Al [19]. In our previous study [11], we observed the formation of an intermediate aluminosilicate phase and finally of the zeolites merlinoite and chabazite. By decreasing the molarity of the alkaline solution from 1 to 0.1 M, we addressed the problem of contemporaneity versus subsequence of the observed processes of diatomite dissolution and formation of new phases and achieved several effects:

1. We decreased diatomite dissolution and let the micropores survive, which yields a hierarchically structured microporous–nanoporous material and allows simultaneous removal of particulate and dissolved matter from aqueous solutions.
2. We completely prevented the formation of zeolites which were very selective but less efficient for heavy metal removal than their amorphous precursors.
3. We produced irregularly shaped particles growing in the micropores of diatoms (Fig. 1). These particles did not yield any peaks of crystalline phases in XRD patterns. However, TEM-SAED patterns reveal that some of these particles are amorphous whereas others are crystalline with  $d$  values of 0.45 and 0.26 nm (Fig. 7). This cannot be explained by a different degree of order, but rather by the small particle size and low percentage of these particles in the sample. TEM-EDX analyses indicate that the chemistry of these particles is similar to that of the intermediate phase in our previous study [11] although they do not act as precursor for zeolite formation in our current experiments. Pore size distribution curves (Fig. 5) indicate that hydrothermal alteration of diatomite using 0.1 M KOH increases both the nanopore (<5 nm) and the micropore (about 100 nm) volume.

Combining the similarity in chemistry and morphology of the observed pore fillings with the intermediate phase mentioned in our previous study [11] with the bimodal pore size distribution allows addressing a hierarchically structured material containing nanosized spherical particles (NSP). The finding of this NSP-containing material is a main outcome of the present study. This material combines micropores of the diatoms with a diameter of about 100 nm with nanopores of the NSP in the range of less than 5 nm according to BJH (Fig. 5). Such hierarchically structured materials are highly promising for technical applications [20]. For example, they might be used for filtration issues removing suspended matter down to 100 nm by the diatom micropores and dissolved ions in the range of few nm by the nanopores of the NSP. For this purpose, it is of great

benefit that the NSP is formed in the pores of the diatoms. This suggests a connectivity of the nanopores of the former with the micropores of the latter which enhances the catalytic activity [6].

Intermediate phases during zeolite formation have been described from experiments without diatoms, e.g., an amorphous phase containing four-membered rings was described [21]. In our previous study [11], we discussed in detail the different Si/Al ratio in hydroxyl-alumino-silicate (HAS) phases. In the present study, measured Si/Al ratios of the NSP are generally higher and K concentrations lower, but are only semi-quantified due to the influence of the diatoms underneath. However, also the NSP produced in this study seem to be of similar chemical composition as the HAS phases observed previously [11]. In contrast to [11], we found some NSP already to be crystalline which is with respect to size (100 nm) and shape (isometric) similar to another study [22], but our observed  $d$  values do neither correspond to those of zeolite P which was observed back then nor to merlinoite which formed later on in our previous experiments [11]. Thus, additional research is needed to identify the crystalline phase occurring in the NSP. In

summary, controlled less-intense dissolution level of diatomite results in precipitation of NSP within still partly remaining diatom skeletons.

### Metal removal experiments

Diatomite altered in alkaline solutions can remove heavy metals from aqueous solutions. Contrarily, it was shown that the capacity of the unaltered diatomite to remove Pb ions from aqueous solution at pH 5 is lacking [11], but relevant at higher pH [23] using diatomite with a similar specific surface area and pore size distribution. The alteration products containing the NSP exhibit a classic ion exchange behaviour defined as metal ion removal with simultaneous K release. This is in agreement with its K content and with other studies showing that in hydrothermal syntheses of various zeolites alkali metals are already present in an amorphous precursor phase [11, 24].

The presence or absence of Al in alteration experiments has a strong impact on the performance of the alteration products in metal ion removal experiments (Fig. 8, 9; Tables 2, 3, 4, 5). When Al was present in alteration

**Table 2** Heavy metal ion removal from the synthetic Cu–Pb–Zn solution (pH 5) after 3 days of reaction with unaltered diatomite ( $t_{\text{syn}} = 0$  day) and diatomite altered for various durations ( $0.25 \text{ day} \leq t_{\text{syn}} \leq 64 \text{ days}$ ) in 0.1 M KOH with 11 mM aqueous Al

$t_{\text{syn}}$ (days)	Cu ( $\mu\text{mol L}^{-1}$ )	Pb ( $\mu\text{mol L}^{-1}$ )	Zn ( $\mu\text{mol L}^{-1}$ )	Si ( $\mu\text{mol L}^{-1}$ )	Al ( $\mu\text{mol L}^{-1}$ )	K ( $\mu\text{mol L}^{-1}$ )
0	417	510	485	n.a.	n.a.	n.a.
0.25	238	36	368	132	2.48	1241
1	223	8	362	133	3.57	1336
2	210	11	369	138	14.91	1421
4	203	9	368	140	1.10	1422
8	183	4	382	149	0.72	1519
16	170	3	382	128	0.91	1533
32	260	10	409	119	1.60	1285
64	241	8	404	125	5.64	1290

Concentrations of dissolved metal ions are given as ratios between the final concentration,  $c_f$ , and the initial concentration,  $c_i$

**Table 3** Heavy metal ion removal from the synthetic Cu–Pb–Zn solution (pH 5) after 3 days of reaction with unaltered diatomite ( $t_{\text{syn}} = 0$  day) and diatomite altered for various durations ( $0.25 \text{ day} \leq t_{\text{syn}} \leq 64 \text{ days}$ ) in 0.1 M KOH without added Al

$t_{\text{syn}}$ (days)	Cu ( $\mu\text{mol L}^{-1}$ )	Pb ( $\mu\text{mol L}^{-1}$ )	Zn ( $\mu\text{mol L}^{-1}$ )	Si ( $\mu\text{mol L}^{-1}$ )	Al ( $\mu\text{mol L}^{-1}$ )	K ( $\mu\text{mol L}^{-1}$ )
0	417	510	485	n.a.	n.a.	n.a.
0.25	331	209	429	159	19.49	525
1	326	206	413	189	17.35	500
2	339	205	435	146	2.35	548
4	288	144	418	182	0.83	729
8	292	164	428	156	1.17	743
16	309	188	444	161	0.40	733
32	315	237	447	155	1.16	592
64	311	318	431	177	3.27	497

Concentrations of dissolved metal ions are given as ratios between the final concentration,  $c_f$ , and the initial concentration,  $c_i$

**Table 4** Heavy metal removal from a mine drainage solution (pH 8) after 3 days of reaction with unaltered diatomite ( $t_{\text{syn}} = 0$  day) and diatomite altered for various durations ( $0.25 \text{ day} \leq t_{\text{syn}} \leq 64 \text{ days}$ ) in 0.1 M KOH with 11 mM aqueous Al

$t_{\text{syn}}$ (days)	Ca ( $\mu\text{mol L}^{-1}$ )	K ( $\mu\text{mol L}^{-1}$ )	Mg ( $\mu\text{mol L}^{-1}$ )	Na ( $\mu\text{mol L}^{-1}$ )	Si ( $\mu\text{mol L}^{-1}$ )	Cd ( $\mu\text{mol L}^{-1}$ )	Pb ( $\mu\text{mol L}^{-1}$ )	Sr ( $\mu\text{mol L}^{-1}$ )	Zn ( $\mu\text{mol L}^{-1}$ )
0	1406	92	742	216	129	0.009	0.029	<0.009	7.67
0.25	779	848	515	140	100	0.001	0.192	0.49	0.14
1	612	930	475	129	241	0.001	0.055	0.40	0.34
2	616	874	492	120	198	0.001	0.112	0.36	0.19
4	628	894	491	116	236	0.002	0.013	0.42	0.09
8	601	924	467	112	279	0.002	0.036	0.47	0.33
16	597	1008	495	120	318	0.002	0.049	0.53	0.35
32	659	696	504	120	230	0.002	0.018	0.46	0.17
64	676	812	481	123	263	0.002	0.218	0.46	0.36

Concentrations of dissolved metal ions are given as ratios between the final concentration,  $c_f$ , and the initial concentration,  $c_i$

**Table 5** Heavy metal removal from a mine drainage solution (0.14  $\mu\text{mol L}^{-1}$  Cd, 0.48  $\mu\text{mol L}^{-1}$  Pb, 17  $\mu\text{mol L}^{-1}$  Zn, pH 8) after 3 days of reaction with unaltered diatomite ( $t_{\text{syn}} = 0$  day) and diatomite altered for various durations ( $0.25 \text{ day} \leq t_{\text{syn}} \leq 64 \text{ days}$ ) in 0.1 M KOH without added Al

$t_{\text{syn}}$ (days)	Ca ( $\mu\text{mol L}^{-1}$ )	K ( $\mu\text{mol L}^{-1}$ )	Mg ( $\mu\text{mol L}^{-1}$ )	Na ( $\mu\text{mol L}^{-1}$ )	Si ( $\mu\text{mol L}^{-1}$ )	Cd ( $\mu\text{mol L}^{-1}$ )	Pb ( $\mu\text{mol L}^{-1}$ )	Sr ( $\mu\text{mol L}^{-1}$ )	Zn ( $\mu\text{mol L}^{-1}$ )
0	1406	92	742	216	129	0.0090	0.029	<0.009	7.67
0.25	784	413	531	172	177	0.0006	0.012	1.31	0.08
1	811	318	533	150	174	0.0007	0.007	1.28	0.08
2	775	326	524	146	201	0.0007	0.005	1.11	0.07
4	724	419	495	141	217	0.0008	0.029	0.96	0.09
8	703	481	520	135	214	0.0007	0.018	1.01	0.10
16	732	479	527	132	212	0.0004	0.013	1.05	0.06
32	740	400	519	128	209	0.0005	0.013	1.09	0.08
64	800	298	508	122	237	0.0005	0.028	1.27	0.14

Concentrations of dissolved metal ions are given as ratios between the final concentration,  $c_f$ , and the initial concentration,  $c_i$

experiments, the Pb ion removal capacity is much higher than for alteration products from experiments without Al. This is by far not as pronounced for Cu and Zn ions. Considering that the NSP are formed more pronounced and are more stable in the presence of Al, it can be concluded that the NSP are responsible for preferential Pb ion removal. This might be explained by a crystalline order of the NSP, which is consistent with TEM-SAED patterns. Cu and Zn ions fit worse into the cavities of these nanocrystals. Hence, the limited fixation of Cu and Zn ions by diatomite altered with 0.1 M KOH can be explained predominantly by ion adsorption. The fact that the more efficient Pb ion fixation by alteration products from experiments with Al correlates with a twice as high release of K supports the suggestion that ion exchange in a cage-like structure takes place, although the Bragg reflections in TEM-SAED patterns could not yet be assigned to a specific crystalline phase.

Unaltered diatomite reduces Pb, Zn and Cd ion concentrations from the natural mine drainage solution (Fig. 9). A previous study with raw diatomite [25] showed maximum adsorption capacities for Zn, Cd and Pb ions of 0.31, 0.18 and 0.10  $\text{mmol g}^{-1}$ , respectively. In this study, 0.5 g diatomite removed only 0.2  $\mu\text{mol Zn}$  ( $=0.4 \mu\text{mol g}^{-1}$ ), 1.25 nmol Cd ( $=2.5 \text{ nmol g}^{-1}$ ) and 0.01  $\mu\text{mol Pb}$  ( $=0.02 \mu\text{mol g}^{-1}$ ) ions which is about one order of magnitude below that value. However, it has to be considered that the obtained final concentrations in the present study are very low and it is obvious that for higher initial concentrations higher amounts of metals could be removed. This is also confirmed by the results with the synthetic solution containing Cu, Pb and Zn ions (Fig. 8).

Alteration of diatomite increases the ion removal capacity with respect to Cd and Zn ions (Fig. 9). As the capacity for Cd and Zn removal is rather independent of the alteration degree and pH, it is suggested that the increase in

metal ion removal capacity is mostly due to the increased surface area. Contrarily, the presence of Al in alteration experiment has an effect, since diatomite altered in pure 0.1 M KOH was more capable of removing Cd, Pb and Zn ions than diatomite altered in Al-containing KOH (Fig. 8). This means that the NSP, which are more “metastable” in presence of Al, performed very well for highly contaminated solutions like the synthetic solution containing Cu, Pb and Zn ions, but decrease metal removal capacities at low metal concentrations like for the used mine drainage solution. It is suggested that adsorption sites of the NSP favour heavy metal ion removal at high metal concentrations but act reversely at low concentrations, whereas the adsorption sites on the diatoms behave vice versa due to their different ion affinities.

The heavy metal removal capacity of altered diatomite is in agreement with our previous study [11], in which we found out that moderately altered diatomite containing the intermediate phase was very effective, but less selective for the removal of Cu, Pb and Zn from synthetic and natural aqueous solutions than intensely altered diatomite containing zeolites. Unfortunately, the NSP-containing material produced in the present study stays far behind both the alteration products obtained in our previous study [11], which contained the intermediate phase and zeolites, with respect to effectivity and selectivity. This can be explained by the fact that the largest percentage of the material applied for metal removal experiments consists still of diatoms, whereas they dissolved more intense in our previous experiments. Consequently, further action is needed to increase the relative percentage of the NSP in the hierarchically structured material.

## Conclusions

Hydrothermal alteration of naturally occurring diatomite in alkaline solution can lead to the formation of crystalline aluminosilicates which has been demonstrated in several studies [11, 20, 22, 26]. Comparingly low molarity of the alkaline solutions used for the present hydrothermal alteration of diatomite suppresses the formation of crystalline aluminosilicates. The herein obtained synthesized product is composed of nanosized spherical particles (NSP) which are partly nanocrystalline and nanoporous and are formed before the complete dissolution of diatoms directly in their microporous skeletons. This enables the application of this tailored material for simultaneous removal of suspended and dissolved contaminants, e.g., in wastewater treatment. Considering the relatively low KOH molarity (0.1 M) of the synthesis solution and the less reaction time required for synthesis (1 day), an economically beneficial technology of diatomite alteration, e.g., by activating packed

columns of diatomite by alkaline solutions will be developed in ongoing studies. Challenging issues for ongoing and future studies comprise of the increase in the relative percentage of the NSP in the hierarchically structured material to adjust distinct metal removal efficiencies without dissolving the diatoms completely to maintain significant microporosity for removal of suspended contaminants.

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## Compliance with ethical standards

**Conflict of interest** The authors declare that they have no conflict of interest.

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